

01 Matter in Our Surroundings (Test - 1)

Time: 45 min

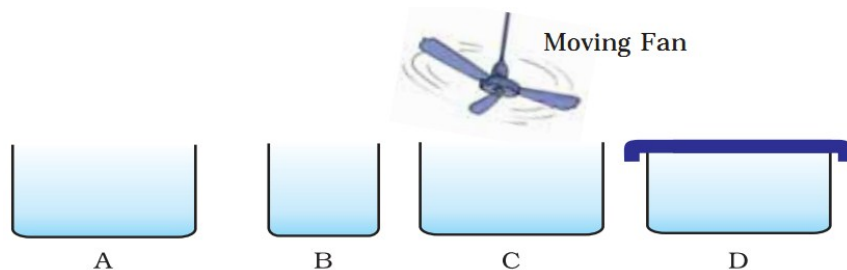
Max Marks: 20

Q. No.	1 to 7	8 to 9
Marks	2	3

1. Fill in the blanks:

- (a) Evaporation of a liquid at room temperature leads to a effect.
- (b) At room temperature the forces of attraction between the particles of solid substances are than those which exist in the gaseous state.
- (c) The arrangement of particles is less ordered in the state. However, there is no order in the state.
- (e) The phenomenon of change of a liquid into the gaseous state at any temperature below its boiling point is called

2. Look at Fig. below and suggest in which of the vessels A, B, C or D the rate of evaporation will be the highest? Explain.



3. Why does the temperature of a substance remain constant during its melting point or boiling point?

4. Give reasons for the following observation:
 The smell of hot sizzling food reaches you several metres away, but to get the smell from cold food you have to go close.
5. Give reasons:
 (a) A gas fills completely the vessel in which it is kept.
 (b) A gas exerts pressure on the walls of the container.
6. Convert the following temperatures:
 (a) 45° C into Kelvin (b) – 158 °C into Kelvin
 (c) 55 K into Celsius (d) 273 K into Celsius.
7. Tabulate the differences in the characteristics of states of matter.
8. It is a hot summer day. Manali and Aarushi are wearing cotton and nylon clothes respectively. Who do you think would be more comfortable and why?
9. Alka was making tea in a kettle. Suddenly she felt intense heat from the puff of steam gushing out of the spout of the kettle. She wondered whether the temperature of the steam was higher than that of the water boiling in the kettle. Comment.